

Durham Research Online

Deposited in DRO:

15 February 2016

Version of attached file:

Accepted Version

Peer-review status of attached file:

Peer-reviewed

Citation for published item:

Collett, C.J. and Massey, R.S. and Maguire, O.R. and Batsanov, A.S. and O'Donoghue, A.C. and Smith, A.D. (2013) 'Mechanistic insights into the triazolydene-catalysed Stetter and benzoin reactions : role of the N-aryl substituent.', *Chemical science.*, 4 (4). pp. 1514-1522.

Further information on publisher's website:

<http://dx.doi.org/10.1039/c2sc22137c>

Publisher's copyright statement:

Use policy

The full-text may be used and/or reproduced, and given to third parties in any format or medium, without prior permission or charge, for personal research or study, educational, or not-for-profit purposes provided that:

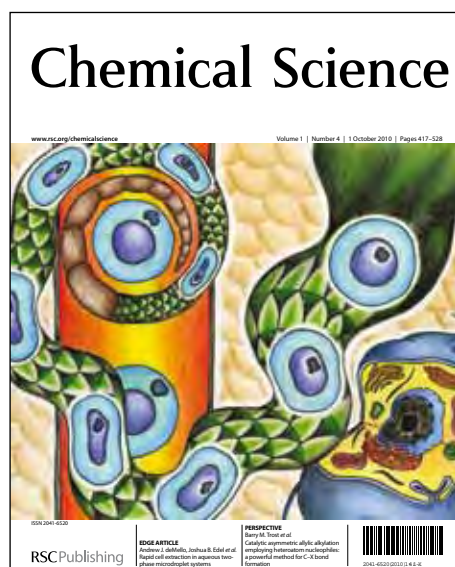
- a full bibliographic reference is made to the original source
- a [link](#) is made to the metadata record in DRO
- the full-text is not changed in any way

The full-text must not be sold in any format or medium without the formal permission of the copyright holders.

Please consult the [full DRO policy](#) for further details.

Chemical Science

Accepted Manuscript



This is an *Accepted Manuscript*, which has been through the RSC Publishing peer review process and has been accepted for publication.

Accepted Manuscripts are published online shortly after acceptance, which is prior to technical editing, formatting and proof reading. This free service from RSC Publishing allows authors to make their results available to the community, in citable form, before publication of the edited article. This *Accepted Manuscript* will be replaced by the edited and formatted *Advance Article* as soon as this is available.

To cite this manuscript please use its permanent Digital Object Identifier (DOI®), which is identical for all formats of publication.

More information about *Accepted Manuscripts* can be found in the [Information for Authors](#).

Please note that technical editing may introduce minor changes to the text and/or graphics contained in the manuscript submitted by the author(s) which may alter content, and that the standard [Terms & Conditions](#) and the [ethical guidelines](#) that apply to the journal are still applicable. In no event shall the RSC be held responsible for any errors or omissions in these *Accepted Manuscript* manuscripts or any consequences arising from the use of any information contained in them.

Cite this: DOI: 10.1039/c0xx00000x

www.rsc.org/chemicalscience

EDGE ARTICLE

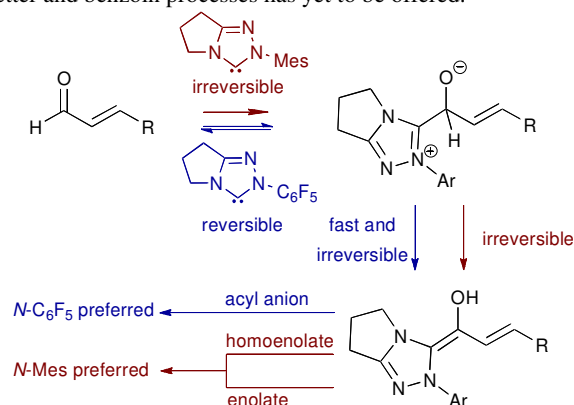
Mechanistic insights into the triazolylidene-catalysed Stetter and Benzoin Reactions: role of the *N*-aryl substituent**Christopher J. Collett, Richard S. Massey, Oliver R. Maguire, Andrei S. Batsanov, AnnMarie C. O'Donoghue* and Andrew D. Smith ***

⁵ Received (in XXX, XXX) Xth XXXXXXXXXX 20XX, Accepted Xth XXXXXXXXXX 20XX
DOI: 10.1039/b000000x

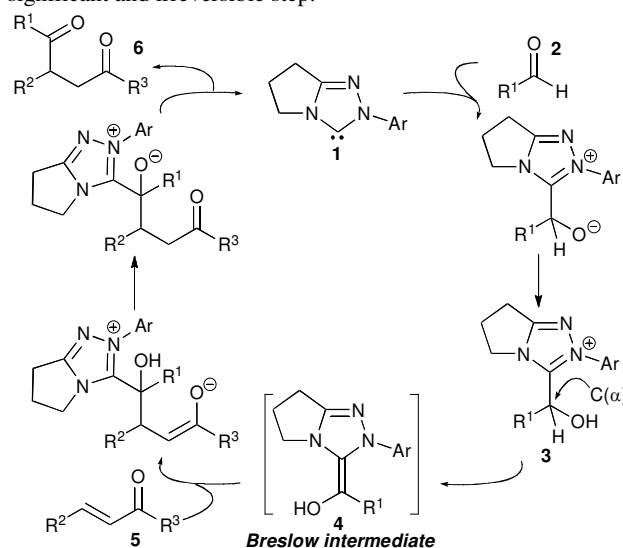
The *in situ* observation, isolation and reversible formation of intermediate 3-(hydroxybenzyl)azolium salts derived from NHC addition to a range of substituted benzaldehydes is probed. Equilibrium constants for the formation of these 3-(hydroxybenzyl)azolium salts, as well as rate constants of hydrogen-deuterium exchange (k_{ex}) at C(α) of these intermediates for a range of *N*-aryl triazolylidenes is reported. These combined studies give insight into the preference of *N*-pentafluorophenyl NHCs to participate in benzoin and Stetter reaction processes.

Introduction

N-Heterocyclic carbenes (NHCs) have been widely employed as organocatalysts,¹ with the triazolylidene molecular class showing remarkable activity in a diverse range of catalytic processes that proceed through acyl anion,² azolium enolate,³ azolium homoenolate,⁴ acyl azolium⁵ or α,β -unsaturated acyl azolium intermediates.⁶ Within the triazolylidene family, the *N*-aryl substituent plays a decisive role in determining catalytic reactivity and selectivity,⁷ with 2,6-substituted *N*-aryl units showing unique reactivity profiles.⁸ For example, *N*-mesityl (*N*-Mes) triazolylidenes are preferred for transformations utilising α -functionalised aldehydes,⁹ while *N*-pentafluorophenyl (*N*-C₆F₅) derivatives usually exhibit increased catalytic activity in Stetter and benzoin processes.¹⁰ Insightful studies from Bode have ascribed the *N*-Mes effect to irreversible addition of the *N*-Mes substituted NHC to the α -functionalised aldehyde, accelerating the formation of the Breslow intermediate (Figure 1).¹¹ To date, a mechanistic rationale for the enhanced performance of electron-deficient *N*-aryl triazolylidenes (*N*-C₆F₅ or *N*-2,4,6Cl₃C₆H₂) in Stetter and benzoin processes has yet to be offered.⁷

Figure 1: Bode's work: The *N*-Mes effect in NHC-mediated processes.

Central to the observed catalytic activity in the benzoin and Stetter reactions is the formation of a common enaminol or Breslow intermediate **4**. Nucleophilic addition of NHC **1** to aldehyde **2** gives tetrahedral intermediate **3**, with deprotonation at C(α) leading to **4**. Onward reaction with an electrophilic Michael acceptor **5**, followed by proton transfer and catalyst regeneration, leads to the product **6** (Scheme 1).¹² While intermediates of the imidazolylidene¹³ and thiazolylidene¹⁴ promoted benzoin reaction similar to **3** have been observed, only limited related studies of triazolylidene-catalysed reactions have been made.¹⁵ Enders and Teles have isolated the 3-(hydroxymethyl)azolium salt addition product of formaldehyde and an NHC,¹⁶ but the concentrations of this intermediate during the reaction have not been monitored. Notably, related structural studies of intermediates in the Stetter reaction have not been established, although mechanistic studies from Rovis indicate that proton transfer from the tetrahedral intermediate **3** is a kinetically significant and irreversible step.¹⁷



Scheme 1: Mechanism of the Stetter reaction.

View Article Online

As part of our studies regarding NHC-catalysed reaction processes,¹⁸ this manuscript describes the *in situ* observation, isolation and reversible formation of intermediate 3-(hydroxybenzyl)azolium salts derived from NHC addition to aldehydes. Equilibrium constants for the formation of these 3-(hydroxybenzyl)azolium salts, as well as rate constants of hydrogen-deuterium exchange (k_{ex}) at C(α) of these intermediates for a range of *N*-aryl triazolinyldenes is reported. These combined studies give insight into the preference of *N*-C₆F₅ NHCs to participate in benzoin and Stetter reaction processes.

In Situ NMR studies: 3-(hydroxybenzyl)azolium salt observation and equilibrium values

To demonstrate the varying catalytic activity of a series of *N*-aryl triazolylidenes, initial *in situ* ¹H NMR spectroscopic studies of the Stetter reaction simply monitored the rate of formation of product **15** from **7** with variation of the *N*-aryl group of the triazolium salt precatalyst (Figure 2).¹⁹ Under catalytic conditions using NEt₃ as a base, electron-deficient *N*-aryl triazolium precatalysts give markedly superior rates of product formation (Ar = C₆F₅ > 2,4,6-Cl₃C₆H₂ > 4-FC₆H₄ > Ph > 4-OMeC₆H₄ > 2,6-OMeC₆H₃ > Mes).²⁰

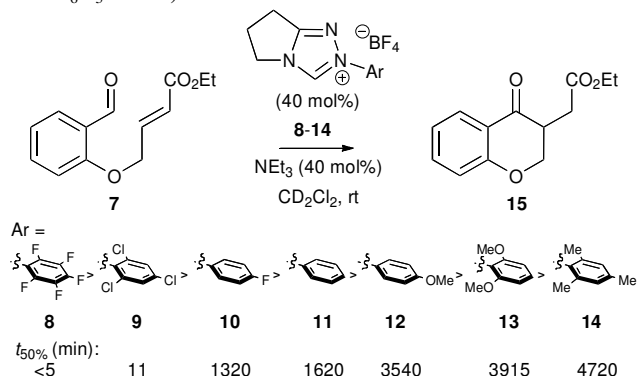


Figure 2: Time taken to reach 50% conversion ($t_{50\%}$) of **7** to **15** with variation of *N*-aryl triazolium substituent.

Further studies utilised ¹H NMR spectroscopic analysis to follow the course of the Stetter reaction of **7** (0.04 M) employing substoichiometric quantities (0.008 M) of *N*-phenyl triazolium precatalyst **11**. Monitoring the reaction in anhydrous CD₂Cl₂ using NEt₃ (0.008 M) as the base showed the initial rapid appearance of an intermediate 3-(hydroxybenzyl)azolium salt **16**, with slower subsequent formation of product **15** over time (Figure 3).²¹ Unambiguous structural determination of **16** was obtained by simply mixing equimolar quantities of triazolium precatalyst **11** with **7** and NEt₃, giving, after 10 minutes, **16** that could be isolated by silica chromatography in 48% yield (Table 1).²² The tetrahedral C(α) geometry within **16** was confirmed by HSQC correlation, complementing the keto tautomer of the Breslow intermediate characterised by NMR spectroscopic analysis by Berkessel and co-workers, who generated this species using a free isolated NHC in THF.²³ Although a small number of related 3-(hydroxybenzyl) thiazolium and imidazolium salts have been prepared by analogous routes,^{13–14} it is notable that **16** can be isolated from synthetically relevant conditions.

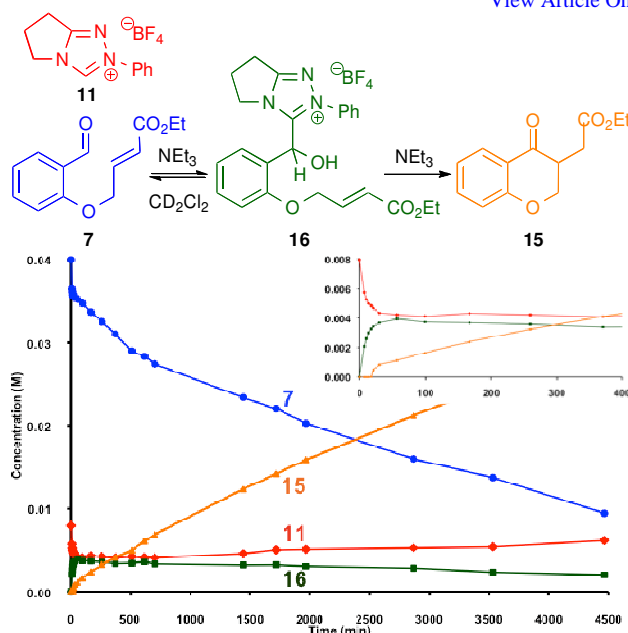


Figure 3: Reaction profile of the Stetter reaction of **7** (0.04 M), catalysed by NHC precursor **11** (0.008 M) with NEt₃ (0.008 M). Inset: expansion of initial time period (<400 min).

Table 1: Preparation and isolation of 3-(hydroxybenzyl)azolium **16**.

Reaction scheme: **7** + **11** (NEt₃ (2 eq), CH₂Cl₂, rt) → **16** (48% yield)

Solvent	δ_{H} C(α)H ^[a]	δ_{C} C(3) ^[a]	δ_{C} C(α) ^[a]
CD ₂ Cl ₂	6.30 (s)	63.2	152.2
DMSO	6.33 d, ($J = 5.4$)	63.2	152.0
CD ₃ OD/DCI	6.36 (s)	65.0	153.7

HSQC correlation (CD₂Cl₂) δ_{H} 6.30 - δ_{C} C(3) 63.2

The stability and reversibility of **16** was then probed using control studies that showed tetrahedral intermediate **16** is stable under either acidic or neutral conditions in CD₂Cl₂. However, treatment of **16** with NEt₃ facilitated rapid equilibration to a mixture of aldehyde **7**, precatalyst **11** and **16**, with relatively slow onwards formation of product **15** (Figure 4).¹⁹ Performing this experiment with **16** (0.008 M) in the presence of additional aldehyde **7** (0.032 M), gave rise to a similar reaction profile to that obtained starting from **7** (0.04 M) and **11** (0.008 M). These observations indicate that **16** is a reversibly formed intermediate from the addition of NHC to the aldehyde, with slower subsequent onwards reaction, presumably through the expected Breslow intermediate, generating product **15**.¹⁹

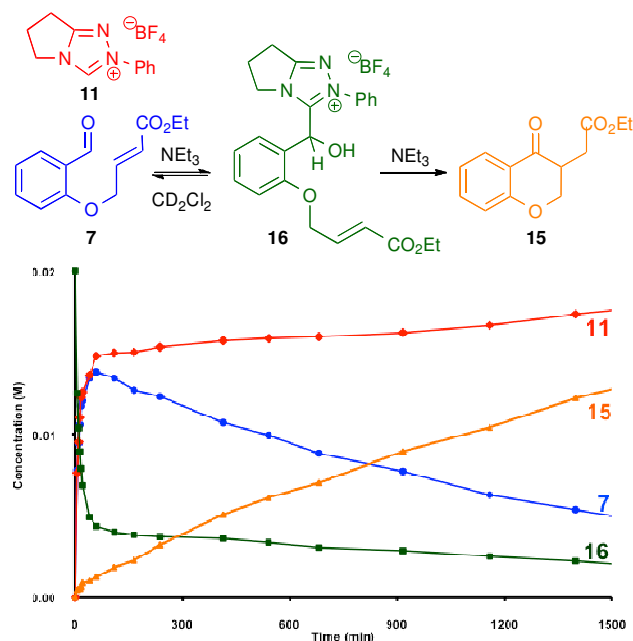


Figure 4: Reaction profile of the reaction of *N*-Ph 3-(hydroxybenzyl)azolum salt **16** (0.02 M) with NEt_3 (0.02 M).

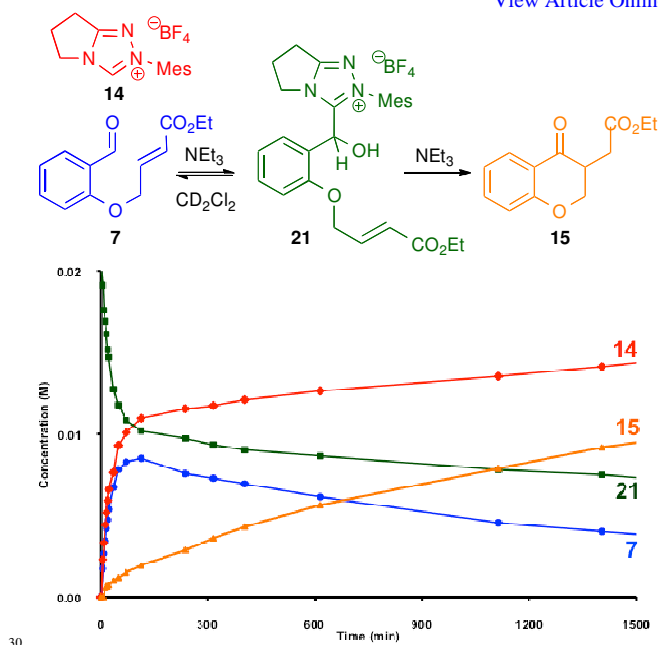
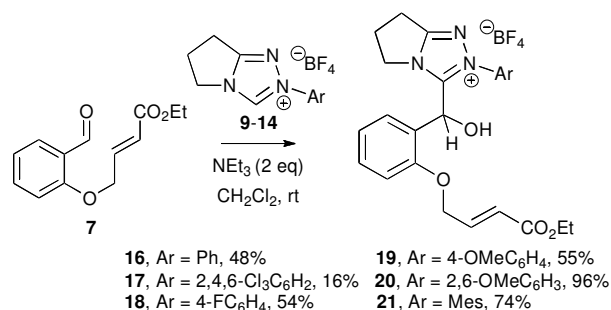


Figure 5: Reaction profile of the reaction of *N*-Mes 3-(hydroxybenzyl)azolum salt **21** (0.02 M) with NEt_3 (0.02 M).

To probe the generality of these studies, the synthesis and isolation of a range of 3-(hydroxybenzyl)azolum salts was investigated. Treatment of a series of *N*-aryl triazolium precatalysts and aldehyde **7** (1:1) with excess NEt_3 gave the corresponding 3-(hydroxybenzyl)azolum salts **17–21** in 16–96% yield (Table 2). Using the *N*- C_6F_5 triazolium precatalyst, the desired tetrahedral intermediate could not be isolated due to rapid conversion into product under these reaction conditions, although electron-deficient *N*-2,4,6- $\text{Cl}_3\text{C}_6\text{H}_2$ and *N*-4- FC_6H_4 analogues **17** and **18** could be isolated and characterised.

Table 2: Preparation of 3-(hydroxybenzyl)azolum salts.



In situ reaction monitoring of the model Stetter transformation using NHC precursors **9–14**, as well as their 3-(hydroxybenzyl)azolum salt products **16–21** was next investigated. In all cases, starting from either the parent azolum salt precatalyst or 3-(hydroxybenzyl)azolum salt, formation of an equilibrium mixture of the corresponding azolum salt, aldehyde **7** and the 3-(hydroxybenzyl)azolum was observed, before relatively slow subsequent onwards reaction to give product **15** (Figure 5, representative example shown using *N*-Mes substituted 3-(hydroxybenzyl)azolum salt **21**).

From the resultant reaction profiles, values for K (the equilibrium constant for the formation of 3-(hydroxybenzyl)azolum salt) were calculated before significant (< 5%) product formation (Table 3).²⁴ These results show that reversible addition is observed in this system even with the *N*-Mes triazolylidene. Notably, significantly larger equilibrium constants are observed using NHC-precatalysts bearing 2,6-substituted *N*-aryl units (*N*-Mes, *N*-2,6- OMeC_6H_3 and *N*-2,4,6- $\text{Cl}_3\text{C}_6\text{H}_2$), whilst electronic variation of 4-substituent leads to minimal perturbation of K .²⁵ Despite the large equilibrium concentration of the *N*-Mes and *N*-2,6- OMeC_6H_3 3-(hydroxybenzyl)azolum salts, the rates of product formation are slower than with NHC precatalysts **8–12**.

Table 3: Equilibrium constants of 3-(hydroxybenzyl)azolum salts.

Ar	K (M^{-1})
2,4,6- $\text{Cl}_3\text{C}_6\text{H}_2$	336
4- FC_6H_4	28
Ph	27
4- OMeC_6H_4	31
2,6- OMeC_6H_3	1283
Mes	140

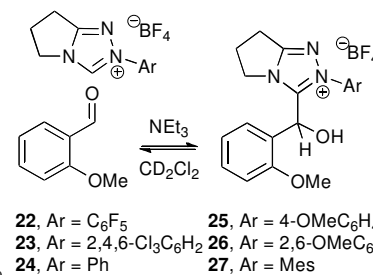
Starting concentrations: **7** 0.04 M, NHC 0.008 M, NEt_3 0.008 M.

Whilst this proved instructive, the effect of the *N*- C_6F_5 substituent upon the equilibrium values could not be evaluated within this system due to rapid onwards reactivity to product. Given the typical recalcitrance of 2-substituted benzaldehydes to participate in homo-benzoin processes,^{10c, 10d} a series of model 3-(hydroxybenzyl)azolum salts **22–27** were prepared from 2-methoxybenzaldehyde, allowing an evaluation of their K values (Table 4).²⁶ Using precatalysts **8**, **9** and **11–14** a similar trend was observed, with 2,6-substituted *N*-aryl NHCs yielding significantly larger K values. The use of *N*- C_6F_5 precatalyst **8** also led to a high K value, suggesting that both steric effects from 2,6- substituents

View Article Online

and electron-withdrawing electronic effects within the *N*-aryl substituent lead to increased *K* values. This increase in *K* reflects the relative stabilities of the respective 3-(hydroxybenzyl)azolium salt in comparison to the starting materials. Simplistically, we assume that 2,6-*N*-aryl substitution within the NHC forces the *N*-aryl ring to adopt an essentially orthogonal orientation with respect to the triazole ring.²⁷ Furthermore, in all 3-(hydroxybenzyl)azolium salts the *N*-aryl group presumably adopts a non-coplanar conformation with the triazole in order to minimize 1,2-steric interactions, with orthogonality enforced with 2,6-*N*-aryl substitution. 3-(Hydroxybenzyl)azolium salt formation may then be favoured for 2,6-*N*-aryl substituted NHCs due to better accommodation of the 3-(hydroxybenzyl) substituent by the *N*-aryl unit. Furthermore, the reverse process may then deviate from expected leaving group ability, being more favourable for non-2,6-*N*-aryl substituted NHCs. Further work will generate additional mechanistic insight by quantifying the rates of the individual processes.

Table 4: Equilibrium constants of 3-(hydroxybenzyl)azolium salts.

		Ar	<i>K</i> (M ⁻¹)
	22, Ar = C ₆ F ₅	C ₆ F ₅	521
	23, Ar = 2,4,6-Cl ₃ C ₆ H ₂	2,4,6-Cl ₃ C ₆ H ₂	601
	24, Ar = Ph	Ph	56
	25, Ar = 4-OMeC ₆ H ₄	4-OMeC ₆ H ₄	46
	26, Ar = 2,6-OMeC ₆ H ₃	2,6-OMeC ₆ H ₃	741
	27, Ar = Mes	Mes	143

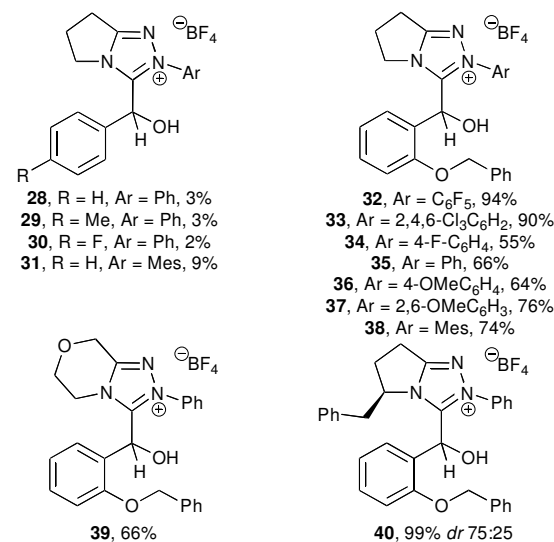
Starting concentrations: 2-methoxybenzaldehyde 0.01 M, NHC 0.002 M, NEt₃ 0.002 M.

Rates constants of exchange (*k*_{ex}) for 3-(hydroxybenzyl)- or 3-(methoxybenzyl)azolium salts

Given the proposed kinetic relevance of the deprotonation step to generate the Breslow intermediate in the Stetter reaction,^{17, 28} the relative rate constants of deuterium exchange (*k*_{ex}, s⁻¹) at the C(α)-H position were investigated by ¹H NMR spectroscopy using either KOD solutions or triethylamine buffers in D₂O/CD₃OD. For Stetter derived 3-(hydroxybenzyl)azolium salts **16–21**, C(α)H exchange could only be monitored alongside competitive dissociation to aldehyde and NHC,²⁹ while attempted *O*-methylation of **16** led to extensive decomposition. To circumvent these issues a series of alternative model 3-(hydroxybenzyl)azolium salts was prepared by triazolinylidene addition to benzaldehyde and a number of substituted benzaldehyde derivatives (Table 5). Only low isolated product yields of 3-(hydroxybenzyl)azolium products were obtained for triazolinylidene addition to either benzaldehyde or 4-substituted benzaldehydes (<10%), with products **28–31** isolated by preparative LC. Despite these poor yields, crystallisation from DCl/D₂O allowed unambiguous structure determination of chloride salt **41** by X-ray crystal structure analysis (Figure 6).³⁰ NHC addition to 2-(benzyloxy)benzaldehyde gave higher isolated yields and showed increased stability to chromatographic

purification, furnishing reasonable to excellent isolated yields (up to 94%) of aldehyde-NHC addition products **32–38**. This synthetic route was extended to morpholine-containing NHC precursors, giving **39**, while a chiral NHC precursor gave **40** as a 75:25 mixture of diastereoisomers in an excellent 99% yield.³¹ 3-(Hydroxybenzyl)azolium salts **28, 32, 33, 35, 37, 38** and **40** were subsequently *O*-methylated³² to facilitate an evaluation of their rates of deuterium exchange.

Table 5: Synthesis of 3-(hydroxybenzyl)azolium salts.



Experimental conditions: NHC precursor (1 equiv), aldehyde (1 equiv), NEt₃ (2 equiv), CH₂Cl₂, rt, 5 min – 1 h.

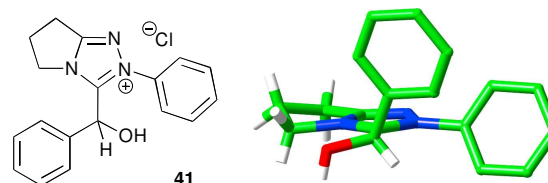


Figure 6. Representation of the X-ray crystal structure of 3-(hydroxybenzyl)azolium chloride **41** (chloride counterion and water of crystallisation not shown for simplicity).

Deuterium exchange studies were carried out upon **36** and **38** and 3-(methoxybenzyl)azolium salts **42–48**. In all cases deuteride catalysed exchange of C(α)-H for deuterium could be monitored without any detritic side reactions (Figure 7).³³ The fastest exchange was observed with electron-deficient *N*-aryl triazolium derivatives (Ar = C₆F₅ > 2,4,6-Cl₃C₆H₂ > Ph > 4-OMeC₆H₄ > Mes > 2,6-OMeC₆H₃, Table 6).³⁴ For **48**, the rate constant of exchange for both diastereoisomers could be evaluated, with the minor diastereoisomer exhibiting an enhanced *k*_{ex} value.³⁵ Within this system, the rate of exchange with variation of *N*-aryl substitution generally parallels the rate of observed product formation in the Stetter transformation.³⁶ This presumably reflects the increased acidity of C(α)-H of the intermediate 3-(hydroxybenzyl)azolium species with increasing electron-withdrawing *N*-aryl substituents, and is consistent with rate determining deprotonation of this species, as postulated by Rovis.¹⁷

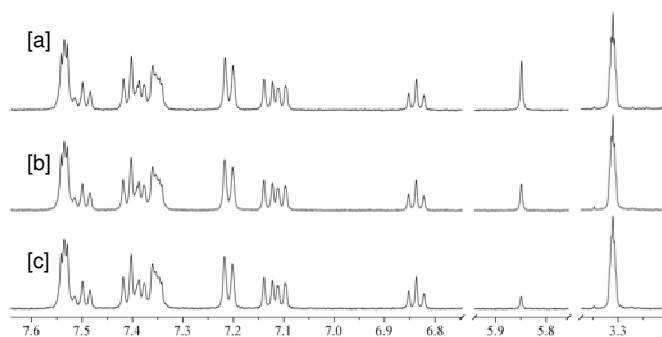
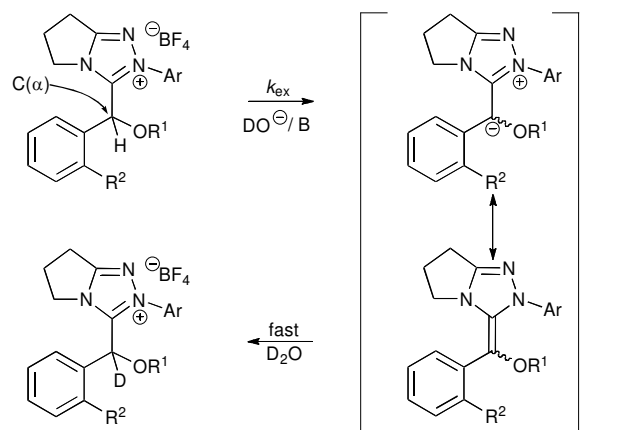


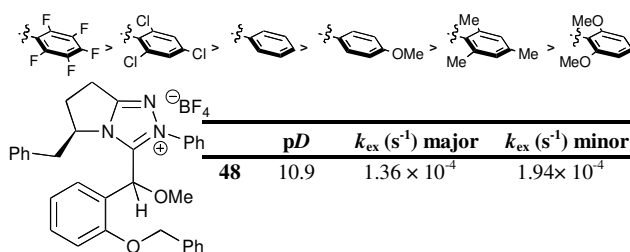
Figure 7. Deuterium exchange of C(α)H (5.85 ppm) relative to CD₃OD (3.31 ppm) of **44** (5 mM) in NEt₃ buffer in 6.5:1 D₂O:CD₃OD at pD 10.9 and 25 °C [a] T = 9 min. [b] T = 109 min. [c] T = 252 min.

Table 6: Relative acidities of 3-(hydroxybenzyl)azolium salts.



	Ar	R ¹	R ²	pD	<i>k</i> _{ex} (s ⁻¹)
36	4-OMeC ₆ H ₄	D	OBn	11.9	6.02 × 10 ⁻⁵
38	Mes	D	OBn	11.9	5.11 × 10 ⁻⁵
42	C ₆ F ₅	Me	OBn	10.9	2.03 × 10 ⁻³
43	2,4,6-Cl ₃ C ₆ H ₂	Me	OBn	10.9	7.30 × 10 ⁻⁴
44	Ph	Me	OBn	10.9	9.60 × 10 ⁻⁵
45	2,6-OMeC ₆ H ₃	Me	OBn	10.9	2.32 × 10 ⁻⁶
46	Mes	Me	OBn	10.9	2.61 × 10 ⁻⁵
47	Ph	Me	H	10.9	3.06 × 10 ⁻⁴

*k*_{ex}: Ar =



Although the formation of the Breslow intermediate is implied by these NMR hydrogen-deuterium exchange studies, direct observation was not possible due to its transient nature in the protic solvent conditions. The direct isolation of Breslow intermediates such as **4** is a widely recognised challenge, although Jordan and co-workers have characterised an *O*-

protected thiazolynylidene-derived enamine using NMR spectroscopy³⁷ and Goldup has reported the observation of an *O*-benzylated intermediate in an imidazolinium-mediated transesterification.³⁸ Most recently, Rovis and co-workers have described the isolation of nitrogen analogues of the Breslow intermediate using 2,6-substituted NHC precursors.³⁹ Mayr has similarly reported the isolation and reactivity of a range of *O*-methylated Breslow intermediates,⁴⁰ whilst Berkessel has reported the isolation of Breslow intermediates derived from free imidazolinium NHCs and benzaldehydes.⁴¹ Despite not being able to observe species such as **4** in these studies, treatment of *O*-methylated derivative **47** with a 10-fold excess of potassium dimsyl in DMSO allowed the acquisition of a UV-Vis spectrum of the enamine with a characteristic λ_{max} at 380 nm and an extinction coefficient (ϵ) of 7846 M⁻¹cm⁻¹, which is consistent with the more extended conjugated system present in **49** relative to the precursor (Figure 8).⁴² Under these conditions, this absorbance decayed to zero over ~20 mins, presumably due to reprotonation of **49** by adventitious water. Under analogous conditions, *O*-methylated derivative **44** exhibited a λ_{max} at 395 nm and an extinction coefficient of 5423 M⁻¹cm⁻¹.

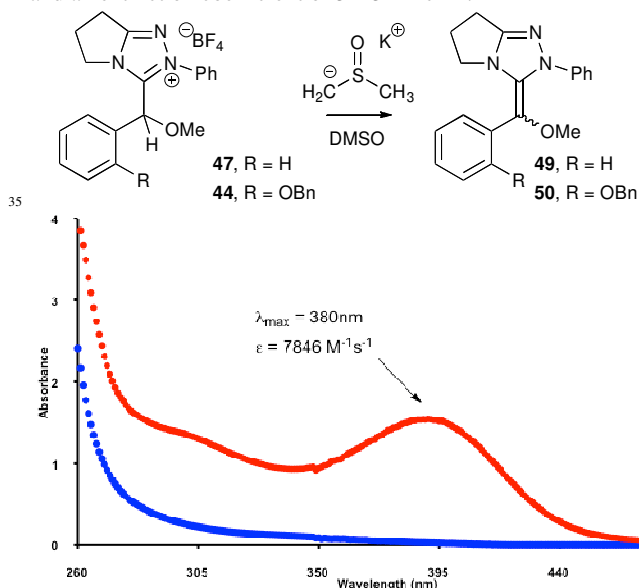


Figure 8. UV-Vis spectra of a 0.2 mM solution of 3-(methoxybenzyl)azolium salt **47** (blue) and **49** (red).

Hydroxybenzylazolium salts as precatalysts

Consistent with these mechanistic studies, the validity of 3-(hydroxybenzyl)azolium salts **18** and **32** as precatalysts for the Stetter reaction was investigated (Figure 9). As expected, treatment of **7** with **18** or **32** (20 mol%) and NEt₃ (20 mol%) in CH₂Cl₂ at rt gave **15** in excellent isolated yield, with the rate of product formation using the electron deficient *N*-C₆F₅ salt **30** significantly faster than the *N*-4-FC₆H₄ salt **18**. Notably, no homo- or crossed-benzoin products were observed in the reaction employing either **18** or **32** as precatalyst.

View Article Online

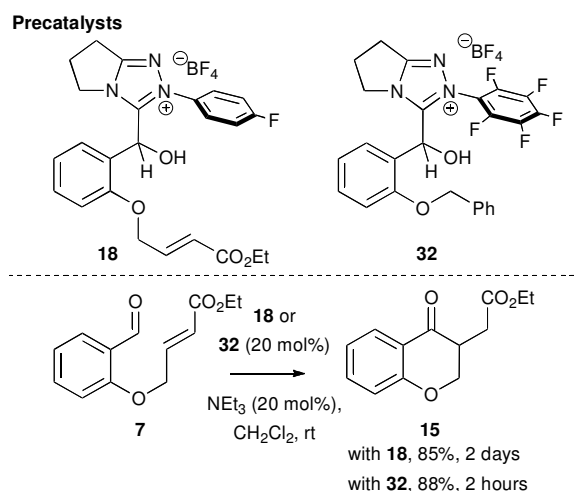


Figure 9. Use of 3-(hydroxybenzyl)azolinium salts as precatalysts in the Stetter reaction.

Similarly, the use of **32** as a precatalyst for the benzoin reaction was probed through performing a crossover reaction with benzaldehyde (Figure 10). Treatment of **32** (20 mol%) with benzaldehyde and NEt₃ (20 mol%) gave preferentially the homo-benzoin product **52**, along with only 6% of the crossed benzoin product **53**. The small yield of the observed crossed product is consistent with current literature, due to the poor reactivity of the 2-substituted aldehyde as an electrophilic component in benzoin type processes.⁴³

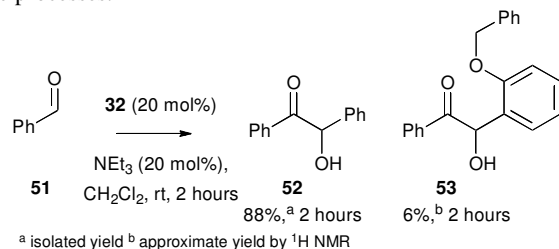


Figure 10. Use of 3-(hydroxybenzyl)azolinium salt **32** as precatalyst in the benzoin condensation.

Conclusions

In conclusion, the isolation and *in situ* spectroscopic observation of aldehyde-NHC addition product intermediates of the Stetter and benzoin reactions is reported. Reaction profiles indicate these intermediates are reversibly formed (irrespective of the *N*-aryl substitution pattern), followed by relatively slow onwards reaction to yield Stetter product. Estimated *K* values suggest that the equilibrium constant is affected by both electronic and steric effects of the *N*-aryl unit, with 2,6- and electron-withdrawing *N*-aryl substitution resulting in significantly larger *K* values. By contrast, electronic variation of the 4-substituent leads to minimal perturbation of *K*. The relative acidities (*k_{ex}*) at C(α) of 3-(methoxybenzyl)azolinium salts provided insight into the effect of catalyst architecture on the key Breslow intermediate-forming step. Despite *N*-2,6-OMeC₆H₃ and *N*-Mes NHC precursors **13** and **14** exhibiting enhanced *K* values (285- and 5-fold vs *N*-Ph congener **11** respectively in the Stetter reaction), slow and presumably rate determining deprotonation¹⁷ (as indicated by *k_{ex}*) inhibits their use as precatalysts in the Stetter reaction. In contrast

N-C₆F₅ and *N*-2,4,6-Cl₃C₆H₂ substituted NHCs benefit from both high *K* values and rapid *k_{ex}* values (21- and 8- fold respectively vs *N*-Ph), rationalising why 2,6-electron withdrawing *N*-substituents are preferred in benzoin and Stetter type processes.⁴⁴ Further work from our laboratories towards developing a full kinetic understanding of these and other NHC-catalysed reaction processes are underway.

Acknowledgements

We thank the Royal Society for a University Research Fellowship (ADS) and the EPSRC (CJC and RM) for funding, and the EPSRC National Mass Spectrometry Service Centre (Swansea). We would also like to thank James Taylor (ADS) and Stuart Leckie (ADS) for their assistance in synthesis of triazolium salts.

Notes and references

- (a) D. Enders, O. Niemeier and A. Henseler, *Chem. Rev.*, 2007, **107**, 5606-5655; (b) N. Marion, S. Diez-Gonzalez and S. P. Nolan, *Angew. Chem. Int. Ed.*, 2007, **46**, 2988-3000; (c) X. Bugaut and F. Glorius, *Chem. Soc. Rev.*, 2012.
- (a) H. Vora and T. Rovis, *Aldrichimica Acta*, 2011, **44**, 3-11; (b) A. T. Biju, N. Kuhl and F. Glorius, *Acc. Chem. Res.*, 2011, **44**, 1182-1195.
- (a) J. Douglas, G. Churchill and A. D. Smith, *Synthesis*, 2012, **44**, 2295-2309; (b) H. U. Vora, P. Wheeler and T. Rovis, *Adv. Synth. Catal.*, 2012, **354**, 1617-1639.
- (a) E. Phillips, A. Chan and K. Scheidt, *Aldrichimica Acta*, 2009, **42**, 55-66; (b) V. Nair, R. S. Menon, A. T. Biju, C. R. Sinu, R. R. Paul, A. Jose and V. Sreekumar, *Chem. Soc. Rev.*, 2011, **40**, 5336-5346.
- (a) N. T. Reynolds, J. Read de Alaniz and T. Rovis, *J. Am. Chem. Soc.*, 2004, **126**, 9518-9519; (b) A. Chan and K. A. Scheidt, *Org. Lett.*, 2005, **7**, 905-908; (c) S. S. Sohn and J. W. Bode, *Org. Lett.*, 2005, **7**, 3873-3876; (d) P.-C. Chiang, Y. Kim and J. W. Bode, *Chem. Commun.*, 2009, 4566-4568; (e) B. E. Maki, E. V. Patterson, C. J. Cramer and K. A. Scheidt, *Org. Lett.*, 2009, **11**, 3942-3945; (f) B. E. Maki, A. Chan and K. A. Scheidt, *Synthesis*, 2008, 1306-1315.
- (a) S. J. Ryan, L. Candish and D. W. Lupton, *J. Am. Chem. Soc.*, 2009, **131**, 14176-14177; (b) J. Mahatthananchai, P. Zheng and J. W. Bode, *Angew. Chem. Int. Ed.*, 2011, **50**, 1673-1677; (c) Z.-Q. Zhu and J.-C. Xiao, *Adv. Synth. Catal.*, 2010, **352**, 2455-2458; (d) S. De Sarkar and A. Studer, *Angew. Chem. Int. Ed.*, 2010, **49**, 9266-9269; (e) J. Guin, S. De Sarkar, S. Grimme and A. Studer, *Angew. Chem. Int. Ed.*, 2008, **47**, 8727-8730; (f) Z.-Q. Rong, M.-Q. Jia and S.-L. You, *Org. Lett.*, 2011, **13**, 4080-4083; (g) Z.-Q. Zhu, X.-L. Zheng, N.-F. Jiang, X. Wan and J.-C. Xiao, *Chem. Commun.*, 2011, **47**, 8670-8672.
- T. Rovis, *Chem. Lett.*, 2008, **37**, 2-7.
- X.-L. Huang, L. He, P.-L. Shao and S. Ye, *Angew. Chem. Int. Ed.*, 2008, **48**, 192-195.
- We thank one of the referees of this manuscript who indicated that some exceptions to this generalisation exist. For example the report by Rovis and co-workers of the synthesis of trans-(γ)-lactams using an *N*-pentafluorophenyl substituted NHC precursor that proceeds via a homoenolate (X. Zhao, D. A. DiRocco and T. Rovis, *J. Am. Chem. Soc.*, 2011, **133**, 12466-12469).

10. (a) D. Enders, K. Breuer and J. H. Teles, *Helv. Chim. Acta*, 1996, **79**, 1217-1221; (b) D. Enders and U. Kallfass, *Angew. Chem. Int. Ed.*, 2002, **41**, 1743-1745; (c) L. Baragwanath, C. A. Rose, K. Zeidler and S. J. Connon, *J. Org. Chem.*, 2009, **74**, 9214-9217; (d) S. E. O'Toole and S. J. Connon, *Org. Biomol. Chem.*, 2009, **7**, 3584-3593; (e) D. Enders, K. Breuer, J. Runsink and J. H. Teles, *Helv. Chim. Acta*, 1996, **79**, 1899-1902; (f) M. S. Kerr, J. Read de Alaniz and T. Rovis, *J. Am. Chem. Soc.*, 2002, **124**, 10298-10299; (g) J. Read de Alaniz and T. Rovis, *J. Am. Chem. Soc.*, 2005, **127**, 6284-6289; (h) D. A. DiRocco and T. Rovis, *J. Am. Chem. Soc.*, 2011, **133**, 10402-10405; (i) D. A. DiRocco, E. L. Noey, K. N. Houk and T. Rovis, *Angew. Chem. Int. Ed.*, 2012, **51**, 2391-2394; (j) T. Jousseau, N. E. Wurz and F. Glorius, *Angew. Chem. Int. Ed.*, 2011, **50**, 1410-1414; (k) D. Enders and J. Han, *Tetrahedron: Asymmetry*, 2008, **19**, 1367-1371; (l) Y. Ma, S. Wei, J. Wu, F. Yang, B. Liu, J. Lan, S. Yang and J. You, *Adv. Synth. Catal.*, 2008, **350**, 2645-2651; (m) I. Piel, M. Steinmetz, K. Hirano, R. Frohlich, S. Grimme and F. Glorius, *Angew. Chem. Int. Ed.*, 2011, **50**, 4983-4987; (n) R. Breslow and E. McNelis, *J. Am. Chem. Soc.*, 1959, **81**, 3080-3082; (o) J. C. Sheehan and T. Hara, *J. Org. Chem.*, 1974, **39**, 1196-1199; (p) W. T. Tagaki, Y. Yano, Y. Bull. Chem. Soc. Jpn., 1980, **53**, 478-480; (q) R. L. Knight and F. J. Leeper, *J. Chem. Soc., Perkin Trans. 1*, 1998, 1891-1894; (r) H. Stetter, *Angew. Chem. Int. Ed.*, 1976, **15**, 639-647.
11. J. Mahatthananchai and J. W. Bode, *Chem. Sci.*, 2012, **3**, 192-197.
12. R. Breslow, *J. Am. Chem. Soc.*, 1958, **80**, 3719-3726.
13. (a) A. Miyashita, A. Kurachir, Y. Matsuoka, N. Tanabe, Y. Suzuki, K.-i. Iwamoto and T. Higashimo, *Heterocycles*, 1997, **44**, 417-426; (b) S. Ohta, S. Hayakawa, K. Nishimura and M. Okamoto, *Chem. Pharm. Bull.*, 1987, **35**, 1058-1069; (c) V. K. Aggarwal, I. Emme and A. Mereu, *Chem. Commun.*, 2002, 1612-1613; (d) L.-W. Xu, Y. Gao, J.-J. Yin, L. Li and C.-G. Xia, *Tetrahedron Lett.*, 2005, **46**, 5317-5320.
14. (a) R. Breslow and R. Kim, *Tetrahedron Lett.*, 1994, **35**, 699-702; (b) Y.-T. Chen, G. L. Barletta, K. Haghighi, J. T. Cheng and F. Jordan, *J. Org. Chem.*, 1994, **59**, 7714-7722; (c) M. J. White and F. J. Leeper, *J. Org. Chem.*, 2001, **66**, 5124-5131; (d) W. Schrader, P. P. Handayani, C. Burstein and F. Glorius, *Chem. Commun.*, 2007, 716-718; (e) M. B. Doughty, G. E. Risinger and S. J. Jungk, *Bioorg. Chem.*, 1987, **15**, 15-30; (f) J. Kiss, R. D'Souza and H. Spiegelberg, *Helv. Chim. Acta*, 1968, **51**, 325-339; (g) J. J. Mielay, G. Bantle, R. G. Votaw, I. A. Rosner and H. Z. Sable, *J. Biol. Chem.*, 1971, **246**, 5213-5219; (h) A. Takamizawa, Y. Hamashima and H. Sato, *J. Org. Chem.*, 1968, **33**, 4038-4045; (i) Y. Kageyama and S. Murata, *J. Org. Chem.*, 2005, **70**, 3140-3147.
15. (a) R. C. Samanta, B. Maji, S. De Sarkar, K. Bergander, R. Frohlich, C. Muck-Lichtenfeld, H. Mayr and A. Studer, *Angew. Chem. Int. Ed.*, 2012, **51**, 5234-5238; (b) K. J. Hawkes and B. F. Yates, *Eur. J. Org. Chem.*, 2008, 5563-5570; (c) J. M. Um, D. A. DiRocco, E. L. Noey, T. Rovis and K. N. Houk, *J. Am. Chem. Soc.*, 2011, **133**, 11249-11254; (d) B. Maji, M. Horn and H. Mayr, *Angew. Chem. Int. Ed.*, 2012, **51**, 6231-6235; (e) G. Doleschall, *Tetrahedron Lett.*, 1980, **21**, 4183-4186; (f) G. Doleschall, *Tetrahedron*, 1976, **32**, 2549-2558; (g) G. Doleschall, *Tetrahedron Lett.*, 1975, **16**, 1889-1892; (h) C. E. I. Knappe, I. I. A. J. Arduengo, H. Jiao, J.-M. Neudörfl and A. Jacobi von Wangelin, *Synthesis*, 2011, **2011**, 3784-3795; (i) C. E. I. Knappe, J. M. Neudörfl and A. J. von Wangelin, *Org. Biomol. Chem.*, 2010, **8**, 1695-1705.
16. J. Henrique Teles, J.-P. Melder, K. Ebel, R. Schneider, E. Gehr, W. Harder, S. Brode, D. Enders, K. Breuer and G. Raabe, *Helv. Chim. Acta*, 1996, **79**, 61-83.
17. J. L. Moore, A. P. Silvestri, J. R. de Alaniz, D. A. DiRocco and T. Rovis, *Org. Lett.*, 2011, **13**, 1742-1745.
18. (a) K. B. Ling and A. D. Smith, *Chem. Commun.*, 2011, **47**, 373-375; (b) C. D. Campbell, C. Concellón and A. D. Smith, *Tetrahedron: Asymmetry*, 2011, **22**, 797-811; (c) N. Duguet, A. Donaldson, S. M. Leckie, E. A. Kallström, C. D. Campbell, P. Shapland, T. B. Brown, A. M. Z. Slawin and A. D. Smith, *Tetrahedron: Asymmetry*, 2010, **21**, 601-616; (d) N. Duguet, A. Donaldson, S. M. Leckie, J. Douglas, P. Shapland, T. B. Brown, G. Churchill, A. M. Z. Slawin and A. D. Smith, *Tetrahedron: Asymmetry*, 2010, **21**, 582-600; (e) J. Douglas, K. B. Ling, C. Concellon, G. Churchill, A. M. Z. Slawin and A. D. Smith, *Eur. J. Org. Chem.*, 2010, **2010**, 5863-5869; (f) N. Duguet, C. D. Campbell, A. M. Z. Slawin and A. D. Smith, *Org. Biomol. Chem.*, 2008, **6**, 1108-1113; (g) J. E. Thomson, C. D. Campbell, C. Concellon, N. Duguet, K. Rix, A. M. Z. Slawin and A. D. Smith, *J. Org. Chem.*, 2008, **73**, 2784-2791; (h) C. Concellón, N. Duguet and A. D. Smith, *Adv. Synth. Catal.*, 2009, **351**, 3001-3009; (i) C. D. Campbell, C. J. Collett, J. E. Thomson, A. M. Z. Slawin and A. D. Smith, *Org. Biomol. Chem.*, 2011, **9**, 4205-4218; (j) R. S. Massey, C. J. Collett, A. G. Lindsay, A. D. Smith and A. C. O'Donoghue, *J. Am. Chem. Soc.*, 2012, **134**, 20421-20432.
19. See SI for further details.
20. Rovis and co-workers have demonstrated that the rate of product formation in the intermolecular Stetter reaction can be enhanced by the addition of bifunctional additives, assumed to assist in the proton transfer step to generate the Breslow intermediate (D. A. DiRocco and T. Rovis, *J. Am. Chem. Soc.*, 2011, **133**, 10402-10405). We chose to examine the reaction without such additives in order the aid the simplicity of analysis.
21. Reactions were typically followed up to 80% conversion to product **15**. See SI for full reaction profiles.
22. One referee correctly questioned whether **16** (or any other hydroxybenzazolum salt in this manuscript) was actually the tetrafluoroborate salt as represented, or the corresponding ylide / alkoxide and Et₃NH⁺. While we are unable to distinguish these possibilities in situ, the product **16** isolated after chromatographic purification contained ¹⁹F NMR resonances at δ_F-153.4 and δ_F-153.5, consistent with it containing a tetrafluoroborate counterion. Similar ¹⁹F NMR data was obtained for **18** and **22** and so for simplicity we represent all compounds as the tetrafluoroborate counterion.
23. A. Berkessel, S. Elfert, K. Etzenbach-Effers and J. H. Teles, *Angew. Chem. Int. Ed.*, 2010, **49**, 7120-7124.
24. Values of *K* were obtained by monitoring the concentrations of the starting materials and 3-(hydroxybenzyl)azolium salt. Agreeable estimates for *K* could be calculated from reaction profiles starting with **7** (0.04 M), NHC (0.008 M) and NEt₃ (0.008 M) or using **7** (0.032 M), 3-(hydroxybenzyl)azolium salt (0.008 M) and NEt₃ (0.008 M). See SI for full details.
25. Glorius postulated that ortho-substituted *N*-aryl substituents would destabilise the initial tetrahedral adduct formed between an NHC and an aromatic aldehyde due to steric effects (See M. Schedler, R. Frohlich, C. G. Daniliuc and F. Glorius, *Eur. J. Org. Chem.*, 2012, 4164-4171; N. E. Wurtz, C. G. Daniliuc and F. Glorius, *Chem. Eur. J.*, 2012 doi 10.1002/chem.201202432). Whilst we believe the enhanced formation of 3-(hydroxybenzyl)azolium salt observed in the case of the *N*-2,6-disubstituted triazolium catalysts results from a more favourable orientation of the *N*-aryl ring (due to the presence of two ortho-substituents), more bulky ortho-substituents do not appear to show this effect. Analogous studies of the benzoin condensation using *N*-Mes and *N*-(2,6-diisopropylphenyl) imidazolium catalysts resulted in the appearance of hydroxybenzyl adduct only in the case of the mesityl catalyst, whilst no reaction was observed using the more sterically hindered 2,6-diisopropylphenyl catalyst (See SI for full details).
26. *K* values were determined as previously described, see SI for full details. During the course of the experiment, none of the crossed benzoin product was observed, presumably due to the poor reactivity of 2-substituted benzaldehydes in the benzoin reaction, consistent with literature studies (see L. Baragwanath, C. A. Rose, K. Zeidler and S. J. Connon, *J. Org. Chem.*, 2009, **74**, 9214-9217; S. E. O'Toole and S. J. Connon, *Org. Biomol. Chem.*, 2009, **7**, 3584-3593).
27. Computational and X-ray crystallographic evidence from Mayr indicates that the *N*-Mes substituent prefers to adopt an almost perpendicular conformation in a range of imidazolium and triazolium carbenes (B. Maji, M. Breugst and H. Mayr, *Angew. Chem. Int. Ed.*, 2011, **50**, 6915-6919). Additionally, a study by Cavallo of NHC ligands concluded that the orientation of the *N*-aryl substituent was determined from the balance between the conjugation energy gained from co-planar geometry and the steric repulsion, with *N*-Ph substituents co-planar, whilst *N*-Mes are almost perpendicular to the heterocycle plane (F. Ragone, A. Poater and L. Cavallo, *J. Am. Chem. Soc.*, 2010, **132**, 4249-4258).

[View Article Online](#)

28. P. Verma, P. A. Patni and R. B. Sunoj, *J. Org. Chem.*, 2011, **76**, 5606-5613.
29. However, the half-life for deuterium incorporation at C(α)H could be estimated, with fastest exchange observed with electron-withdrawing *N*-aryl units (Ar = 4-FC₆H₄ > Ph > 4-OMeC₆H₄ > Mes). See SI for further details.
30. **41** and the dichloride salts of **29** and **30** (**54** and **55** respectively) were characterised by single-crystal X-ray diffraction of their D₂O monosolvates which are mutually isotopic (isomorphous). Crystallographic data is available free of charge from the Cambridge Crystallographic Data Centre, www.ccdc.ac.uk/data-request/cif, as CCDC-895406 (**41**), 895407 (**54**) and 895408 (**55**).
31. Treatment of the isolated major diastereoisomer to the reaction conditions led to a 75:25 mix of diastereoisomers, suggesting the obtained *dr* is a result of a thermodynamic equilibrium.
32. 3-(Methoxybenzyl)azolium salts **42-48** were synthesised from methylation of **28**, **32**, **33**, **35**, **37**, **38** and **40** using diazomethane or TMS-diazomethane. See SI for further details. For these compounds, exchange could be monitored without dissociation or decomposition.
33. Experiments were carried out at a substrate concentration of 5 mM at 25 °C and *I*=1 (KCl) in 6.5:1 D₂O/CD₃OD. Exchange was monitored for two half lives and a semi-logarithmic plot of the fraction of unexchanged substrate (*f*_s) against time gave *k*_{ex}, the pseudo-first order rate of exchange at that *pD*. In the case of **36** and **38**, measurements at other *pD* values were not possible due to competing dissociation to aldehyde and NHC at lower *pD* values and ring opening at higher *pD* values. See SI for full details.
34. It is assumed that deprotonation by deuteroxide is significantly faster than deuteration of the Breslow intermediate in all cases so that *k*_{ex} reflects the rate constant for formation of the solvent equilibrated enaminol/enamine. Using stopped flow spectrophotometry Jordan has measured rate constants for the reprotonation of *O*-methylated thiazolylidene-derived enamine derivatives in the range of 300-540 s⁻¹ (G. L. Barletta, Y. Zou, W. P. Huskey and F. Jordan, *J. Am. Chem. Soc.*, 1997, **119**, 2356-2362), which are 10⁷-fold higher than corresponding rate constants for formation of the intermediate at *pH* 10.5 (~*pD* 10.9).
35. Upon completion of the deuterium exchange experiment of **48** the C(α)-deuterated **48** formed exhibited a decreased *dr*, suggesting enamine deuteration is not completely stereoselective.
36. The only exception to this is the 2,6-OMeC₆H₃ substituted analogue **45**, which exhibits a lower *k*_{ex} value than **46**, despite displaying a faster rate of product formation in the Stetter reaction. We presume that the greatly enhanced *K* value (Table 3) compensates for the slower rate of deprotonation, giving a faster overall rate of reaction. Jordan has reported a second order rate constant, *k*_{HO} = 0.019 M⁻¹s⁻¹ for hydroxide ion deprotonation of the 2-(methoxyphenylmethyl)-3,4-dimethylthiazolium salt to the corresponding enamine using stopped-flow spectrophotometry (G. L. Barletta, Y. Zou, W. P. Huskey and F. Jordan, *J. Am. Chem. Soc.*, 1997, **119**, 2356-2362). Using this value, we can calculate a first order rate constant for deprotonation by hydroxide at *pH* 10.5 (~*pD* 10.9) of 8 x 10⁻⁶ s⁻¹, which is 3-250-fold smaller than any of the *k*_{ex} values in Table 6 at an equivalent *pD*.
37. F. Jordan, Z. H. Kudzin and C. B. Rios, *J. Am. Chem. Soc.*, 1987, **109**, 4415-4416.
38. L. Pignataro, T. Papalia, A. M. Z. Slawin and S. M. Goldup, *Org. Lett.*, 2009, **11**, 1643-1646.
39. D. A. DiRocco, K. M. Oberg and T. Rovis, *J. Am. Chem. Soc.*, 2012, **134**, 6143-6145.
40. B. Maji and H. Mayr, *Angew. Chem. Int. Ed.*, 2012, **51**, 10408-10412.
41. A. Berkessel, S. Elfert, V. R. Yatham, J. M. Neudorfl, N. E. Schlörer and J. H. Teles, *Angew. Chem. Int. Ed.*, 2012, **51**, 12370-12374.
42. This value is consistent with the reported extinction coefficient of 14000 M⁻¹cm⁻¹ at 410 nm in dioxane/DMSO for the enamine generated from a 2-(methoxyphenylmethyl)-3,4-dimethylthiazolium salt (G. Barletta, W. P. Huskey and F. Jordan, *J. Am. Chem. Soc.*, 1992, **114**, 7607-7608).
43. None of the homo-benzoin product formed between 2-(benzyloxy)benzaldehyde was observed over the course of the reaction, consistent with current literature (see S. E. O'Toole, C. A. Rose, S. Gundala, K. Zeitler and S. J. Connon, *J. Org. Chem.*, 2010, **76**, 347-357; M. Y. Jin, S. M. Kim, H. Han, D. H. Ryu and J. W. Yang, *Org. Lett.*, 2011, **13**, 880-883).
44. We are grateful to a referee who pointed out that the electronic determinant of reaction rate in the intramolecular Stetter transformation (superior rates with electron-deficient *N*-aryl triazolium precatalysts) is in contrast to many other classes of NHC-catalysed reactions. For example, typical reactions that proceed via azolium enolates appear to be dominated by steric constraints. For a computational study see: S. E. Allen, J. Mahatthananchai, J. W. Bode and M. C. Kozlowski, *J. Am. Chem. Soc.*, 2012, **134**, 12098-12103.